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A Simple and Cost-efficient Technique to Generate Hyperpolarized Long-lived ^{15}N-^{15}N Nuclear Spin Order in a Diazine by Signal Amplification by Reversible Exchange

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ABSTRACT

Signal Amplification by Reversible Exchange (SABRE) is an inexpensive and simple hyperpolarization technique that is capable of boosting Nuclear Magnetic Resonance (NMR) sensitivity by several orders of magnitude. It utilizes the reversible binding of *para*-hydrogen, as hydride ligands, and a substrate of interest to a metal catalyst to allow polarization transfer from *para*-hydrogen into substrate nuclear spins. While the resulting nuclear spin populations can be dramatically larger than those normally created, their lifetime sets a strict upper limit on the experimental timeframe. Consequently, short nuclear spin lifetimes are a challenge for hyperpolarized metabolic imaging. In this report we demonstrate how both hyperpolarization and long nuclear spin lifetime can be simultaneously achieved in nitrogen-15 containing derivatives of pyridazine and phthalazine by SABRE. These substrates were chosen to reflect two distinct classes of $^{15}N_2$-coupled species that differ according to their chemical symmetry and thereby achieve different nuclear spin lifetimes. The pyridazine derivative proves to exhibit a signal lifetime of ca. 2.5 minutes and can be produced with a

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signal enhancement of *ca*. 2,700. In contrast, while the phthalazine derivative yields a superior 15,000-fold ^{15}N signal enhancement at 11.7 T it has a much shorter signal lifetime.

I. INTRODUCTION

Despite the many significant advances that have taken place in Nuclear Magnetic Resonance (NMR) since its inception, poor sensitivity still limits full utility. This low sensitivity arises because NMR relies on the Boltzmann distribution to create population imbalances between the nuclear spin orientations it probes.[1] Whilst ^{1}H detection offers maximum sensitivity, the signal amplitude still originates from a difference of just 1 in each 32,000 ^{1}H spins at room temperature within a 9.4 T magnet.[1] This problem is even more pronounced for low-γ nuclei such as ^{13}C and ^{15}N, where in the latter case just 1 in every 300,000 ^{15}N nuclear spins contribute positively at this field.[1, 2]

Recent developments in hyperpolarization techniques that improve sensitivity have allowed the development of magnetic resonance applications that were previously thought to be beyond the techniques reach.[3] This builds from the fact that techniques such as Dynamic Nuclear Polarization (DNP)[4] and Spin Exchange Optical Pumping (SEOP)[5] provide unprecedented levels of signal enhancement for carbon-13, nitrogen-15 and xenon-129 spin detection. While these developments have been applied to *in vivo* study,[6-8] they often involve high-cost instrumentation[4] which acts to restrict their utilization.

An alternative approach involving *para*-hydrogen (*p*-H_2) as a source of polarization is gaining popularity due to its speed and simplicity.[9, 10] Methods involving *p*-H_2 are referred to as *Para*-Hydrogen Induced Polarization (PHIP) approaches and classically use a metal catalyst to add *p*-H_2 to an unsaturated substrate *via* a hydrogenation step. However, a variant of PHIP called Signal Amplification by Reversible Exchange (SABRE) has greatly expanded the remit of the PHIP method as it does not induce chemical change in the substrate.[11] SABRE instead employs reversible substrate and *p*-H_2 binding to a catalyst to transfer polarization from the *p*-H_2 derived hydride ligands into a

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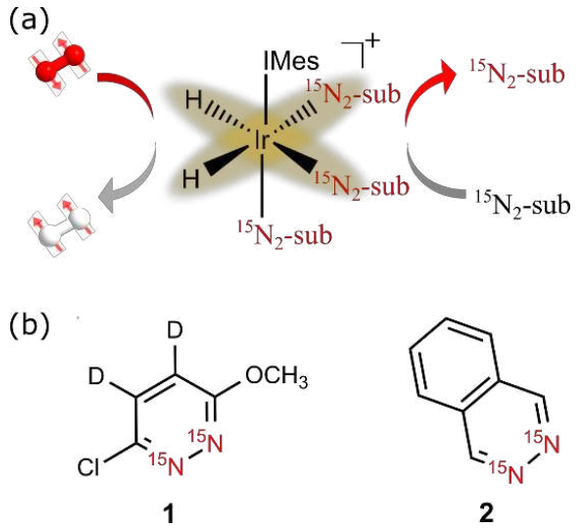
selected substrate under appropriate resonance conditions (Scheme 1).[12] Since its inception, SABRE has become successful at hyperpolarizing a growing range of important materials such as nicotinamide, methyl nicotinate, imidazole, diazirines, metronidazole, amines and pyruvate.[13-20]



SCHEME 1: (a) Schematic depiction of the SABRE hyperpolarization method; *p*-H_2 and substrate (sub) bind reversibly to an iridium catalyst to induce polarization transfer. (b) Structures of the substrates used in this study—3-chloro-6-methoxy-4,5-d_2-pyridazine-$^{15}N_2$ (**1**) and phthlazine-$^{15}N_2$ (**2**).

The hyperpolarization of heteronuclei provides two crucial advantages over normal ^{1}H magnetic resonance imaging (MRI) – (a) an essentially background-free signal and (b) potentially long magnetic state lifetimes. This is reflected in the fact that the greatest success of DNP to date has been the hyperpolarization of ^{13}C nuclei in isotopically labelled pyruvate for the subsequent study of metabolic pathways linked to cancer.[6, 21-23] Hyperpolarized ^{15}N offers similar advantages to ^{13}C detection and the feasibility of its use *in vivo* has been established previously for ^{15}N-choline.[24] As the relative molar receptivity of ^{15}N is just 1.04×10^{-3} and ^{13}C 1.59×10^{-2} with respect to ^{1}H, the use of hyperpolarization is critical for such heteronuclei detection.[1]

Warren and co-workers have demonstrated that ^{15}N targets can be produced with high levels of hyperpolarization together with long magnetic state lifetimes using a variation of SABRE they termed

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SABRE-SHEATH.[25-29] It simply uses a mu-metal shield to enable efficient and direct polarization transfer from the hydride ligands of the catalyst to heteronuclei. A significant breakthrough was reflected in their studies of diazirines which were found to display both longitudinal magnetization and long lived singlet states after polarization transfer.[16] ^{15}N polarization levels of ca. 5% were reported and the associated singlet state had a lifetime of 23 min. This singlet state was revealed by the use of chemical asymmetry and built from work by Levitt and co-workers who illustrated how long-lived singlet states (LLS) sustain nuclear spin lifetimes beyond those of the normal T_1 timescale through storage in disconnected Eigen states[30, 31] that are immune to the major mechanisms of relaxation.[32] Examples of such systems have been found where these long-lived states have lifetimes that exceed 1 hour, or 50 times the more usual T_1 timescale, in room temperature solution.[33, 34]

In this work, we use the SABRE variant SABRE-SHEATH to hyperpolarize two $^{15}N_2$-based diazines and rationalize the basis of a simple route to their detection over long-time-scales. To broaden applicability, these agents were selected to represent two kinds of substrate that differ according to whether their coupled ^{15}N-spins are chemically or magnetically different.

II. EXPERIMENTAL METHODS AND RESULTS

The *p*-H_2 used in this SABRE hyperpolarization study was created in more than 92% purity using an in-house *para*-hydrogen generator.[35] Samples were prepared by mixing 5 mM of [IrCl(COD)(IMes)] (IMes = 1,3-bis(2,4,6-trimethylphenyl)imidazol-2-ylidene) with 30 mM of the substrate (**1** or **2** of Scheme 1) in 0.6 mL of methanol-d_4 in a 5 mm NMR tube fitted with a J. Young's Tap. After degassing using a freeze-pump-thaw method, the samples were activated by the introduction of H_2 at a pressure of 3 bar. SABRE hyperpolarization experiments were then completed by filling the NMR tubes with *p*-H_2 (3 bar) and subsequently shaking them vigorously in the specified magnetic field before detecting the resulting signal inside a high field NMR spectrometer (11.75 T). In these experiments, a mu-metal shield was used to reduce the background magnetic field to around 1000

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times its normal value so that a mG top-up field can be applied to the sample through the application of a solenoid.[17] The SABRE/SABRE-SHEATH hyperpolarization and sample transfer steps take place over 10 to 20 seconds. [36, 37] Since SABRE is reversible, sample re-hyperpolarisation can be achieved within just a few seconds by repeating this procedure with fresh *p*-H_2. In this way, accuracy and relaxation effects can readily be probed. The NMR measurements that feature in the final observation step were carried out at 298 K on an 11.75 T Bruker Avance III spectrometer using a TBI probe.

Pyridazine **1** contains a pair of coupled ^{15}N spins that are chemically different. Earlier studies of several related pyridazine based substrates confirmed they can provide access to good ^{1}H-SABRE hyperpolarization levels, thereby indicating the suitability of these systems.[28, 29, 38] The hyperpolarization of 3,6-dichloropyridazine-$^{15}N_2$ has also previously been reported.[39] The pyridazine motif is itself prevalent in a range of pharmacologically active agents and hence screening their NMR detection and magnetic state lifetimes is sensible.[40]

The chemical shift between the two inequivalent ^{15}N sites in **1** was quantified to be 29.3 ppm (1485 Hz at 11.75 T), with a mutual spin-spin coupling of 23 Hz connecting them. During the SABRE process, **1** and a pair of *p*-H_2 derived hydride ligands bind to the iridium catalyst ([Ir(H)$_2$(NHC)(**1**)$_3$]Cl) to temporarily create an AA'BC type 4-spin system at low-magnetic field where the *trans* hydride-^{15}N coupling is around 20 Hz, the hydride-hydride coupling ~-8 Hz and the retained ^{15}N-^{15}N ~|20| Hz. Consequently, the SABRE transfer mechanism for diazines **1** and **2** lead to direct population of the corresponding $^{15}N_2$-spin system singlet state after dissociation.[16, 18, 27, 38, 41] This is the result of the fact that the J_{HH} and J_{NN} couplings are sufficiently close in size to enable the ^{15}N-singlet to become populated in fields where the difference in chemical shift between the two bound nitrogen resonances is smaller than the ^{15}N-^{15}N *J*-coupling.

Once SABRE hyperpolarization experiments were performed according to the aforementioned protocol in the case of **1** adiabatic transfer to 11.75 T enables the observation of two ^{15}N NMR

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signals, as detailed in Figure 1a, after a 90˚ hard observation pulse. These signals possess an 'up-up-down-down' pattern that indirectly confirms the creation of ^{15}N-singlet spin character in **1** after completion of SABRE.[17, 42, 43] This is the result of probing a high-field state of the form $I_zS_z + I_z - S_z$ which leads to two observable doublets, of opposite relative phase, when interrogated by a 90˚ read pulse. When the same process was repeated, but a 9˚ flip angle used, the resulting NMR spectrum yields detectable outer-line transitions as shown in Figure 1b which further confirm the presence of initial singlet spin character as the I_zS_z term which leads to a pair of antiphase doublets now adds to the earlier signal. These observations also show that the resulting state does not decohere rapidly which confirms that the presence of the methyl substituent has minimal effect on the signals lifetime.[44, 45] This is in agreement with the failure to observe any scalar coupling between the methyl groups protons and the ^{15}N centers. For comparison purposes, Figure 1c shows the corresponding thermally polarized ^{15}N NMR spectrum that was acquired in conjunction with signal averaging over 1000 scans where the delay between measurements is 120 sec. Consequently, this control measurement took over 33 hours to make. On the basis of these data, a signal enhancement factor of 1250 could be determined at 11.75 T, relative to the thermally polarized NMR spectrum. These measurements were repeated using different polarization transfer field values in the range 1 to 10 mG and little intensity variation was observed which is expected for a direct singlet transfer pathway.

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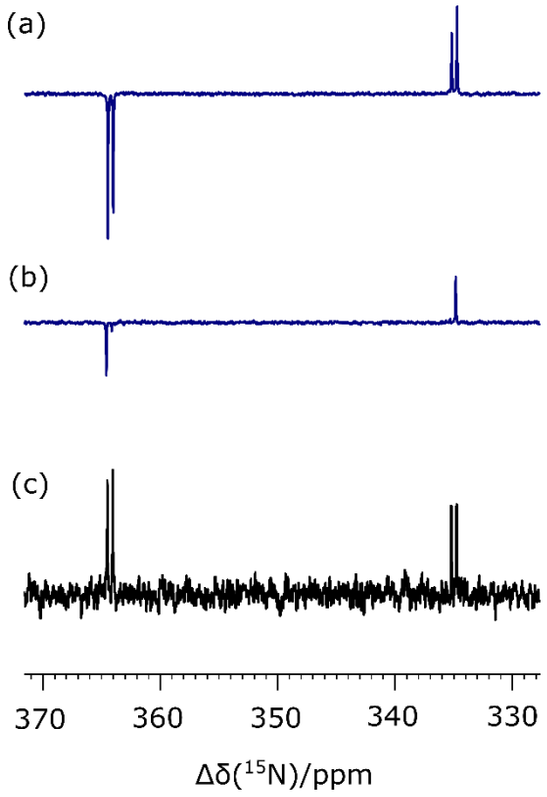

FIGURE 1: ^{15}N NMR spectra associated with **1**: (a) single-shot hyperpolarized SABRE-SHEATH experiment detected at 11.75 T by a 90˚pulse, (b) detected by a 9˚pulse and (c) the corresponding ^{15}N NMR spectrum after 1000 averages.

Since the SABRE signal that is created in these experiments originates in the corresponding singlet state, it's lifetime should be much longer than that associated with more usual T_1 decay. Furthermore, since chemical shift anisotropy (CSA) is the major source of singlet order relaxation, this period should be extended significantly with lower magnetic field storage.[46] This situation is complicated by the fact that the singlet (S_0) state of the free material connects directly with the shorter lived triplet (T_o) state which will act to reduce its population and therefore the high-field lifetime. However, when the substrate is bound it will exhibit an even more reduced S_0 lifetime due to the potential to transfer polarization into the hydride ligands in the reverse of the initial ^{15}N polarization transfer step and the existence of spin-spin couplings to the hydride ligands can also

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lead to the creation of triplet derived magnetization. These effects can be readily evaluated though by changing the metal concentration.

We therefore first measured the effective lifetime of the magnetization created under SABRE after storage in three magnetic fields. For the high magnetic field value, we used the 11.75 T field of the NMR system and determined the signal lifetime to be 35.8 ± 5.8 s. Next, the sample was examined after storage at 0.3 T and the lifetime of the signal increased to 56.5 ± 12.6. The 0.3 T field was selected because of the work of Shchepin *et* al. where they found it proved suitable for hyperpolarized T_1 extension.[47] Upon storage in the mu-metal shield, the signal lifetime became 118.3 ± 20.4 s. The normalized signal intensities used in obtaining these values alongside the corresponding exponential fits are shown in Figure 2 and the results are detailed in Table 1. As indicated above, these signal lifetimes are each measured in the presence of the active SABRE catalyst. They are therefore further compressed by the reversible interaction of **1** with the catalyst which more efficiently breaks the symmetry of the magnetic state during the ligation event as $\delta\Delta$ increases to ~3000 Hz for the ligand bound *trans* to hydride at 11.75 T. Consequently, when these measurements are repeated with a 50-fold excess of **1** based on iridium, rather than the 6-fold described first, these lifetimes are extended. Now, the signal lifetime becomes 48.8 ± 7.1 s at 11.75 T whilst at zero field (mu-metal shield) it became 155.5 ± 15.4 s. Hence, we can conclude that the lifetimes can be substantially improved in the presence of a larger excess of substrate which reduces the propensity for magnetization decay through ligand exchange. However, a significant drop in signal enhancement factor to *ca*. 200-fold is also observed at this higher substrate loading and therefore a balance between signal size and lifetime needs to be considered based on the desired application.

Due to the long lifetime of the created hyperpolarized ^{15}N signal of **1**, we could further improve the signal enhancement achieved with a sample containing 5 mM [IrCl(COD)(IMes)] and 30 mM of **1** by extending the polarization transfer times. When a 25 second polarization time was employed the

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visible signal gain increased significantly to 2700-fold. This represents the detection of a signal that is twice as large as that achieved with a 10 second transfer time. However, when the polarization time was increased above 30 seconds, the ^{15}N signal gain decreased to 2500-fold which reflects the finite volume of *p*-H_2 that is present in the sealed NMR tubes used in this study.

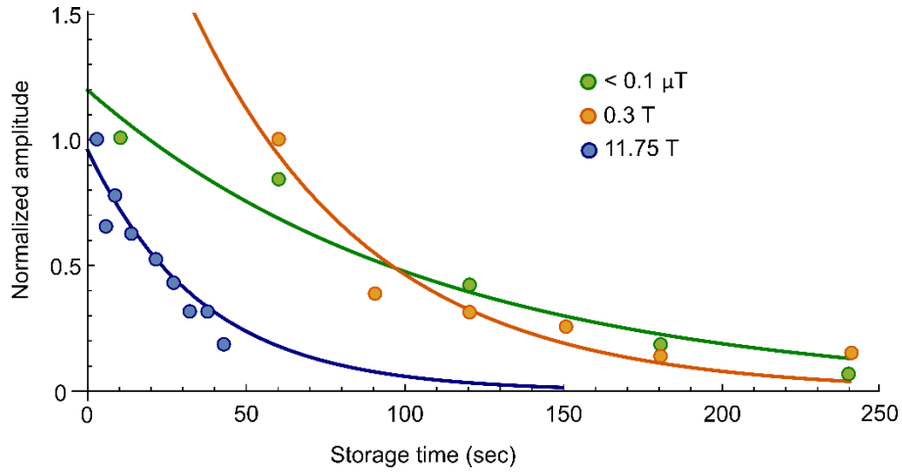

FIGURE 2: Normalized signal amplitudes of the ^{15}N hyperpolarised NMR signals seen for **1** (circles) after SABRE-SHEATH as a function of sample storage time. Data points are fitted to an exponential (solid curves) which yields the signal lifetimes reported in Table 1 for a precatalyst to **1** loading of 1:6. Sample storage took place at 11.75 T (blue), 0.3 T (orange) and 0 T (green).

TABLE 1: ^{15}N polarization levels (enhancement factor and %), and signal lifetimes for **1** and **2** at the specified storage fields achieved with the precatalyst [IrCl(COD)(IMes)]. All measurements were made at 11.75 T and 298 K.

Agent	Enhancement factor (ε) and net polarization (P%)	Hyperpolarised signal lifetime at 11.75 T	Hyperpolarised signal lifetime at 0.3 T	Hyperpolarised signal lifetime at 0.1 μT
1	ε: 1250 P: 0.5%	T_{LLS}: 35. 8 ± 5.8 s	T_{LLS}: 56.5 ± 12.6 s	T_{LLS}: 118.3 ± 20.4 s

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2	ε: 4800 P: 2%	T_1: 5. 8 ± 0.2 s	T_1: 56.0 ± 2.5 s	T_1: 21.0 ± 6.3 s

The second substrate, phthalazine **2**, has a chemically equivalent but magnetically distinct $^{15}N_2$-spin system due to the associated ^1H couplings. It was probed under SABRE-SHEATH conditions inside a mu-metal shield as described above.[17, 48] The presence of the α-proton substituents on the ring system, and their visible couplings to ^{15}N, will enable decoherence of any S_o term that is created through SABRE and thereby make the resulting states visible to NMR.[25, 46] However, as indicted earlier, the transient binding of **2** to a metal complex will break both the chemical and magnetic symmetry of this ^{15}N pair, thereby providing not only a route to see both bound and free material, but also a route to further decohere the singlet state, in a process whose effect will again be concentration dependent.

For these measurements we initially employed a solution containing 5 mM of [IrCl(COD)(IMes)] and 30 mM of **2**. Figure 3 shows the resulting series of hyperpolarized ^{15}N NMR signals for **2** that were observed after the application of a 90 degree observation pulse as a function of the polarization transfer field strength and we see the signal reaches maximum amplitude at 4.5 mG.

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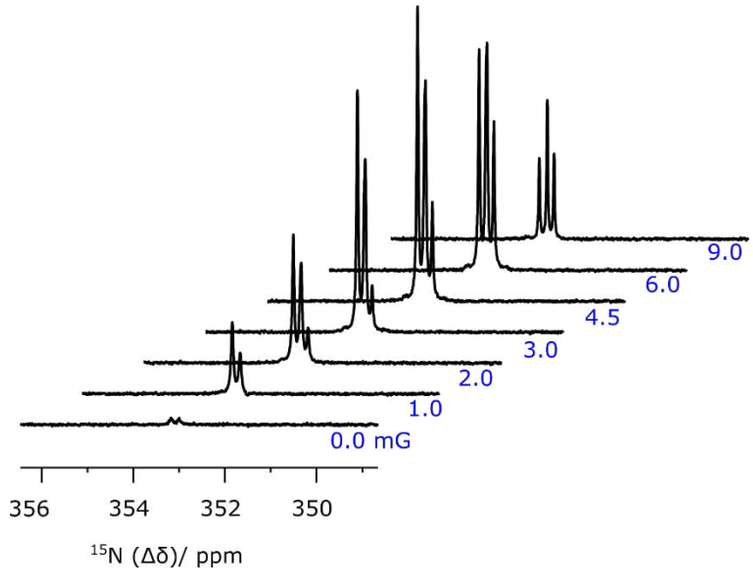

FIGURE 3: Hyperpolarized ^{15}N NMR spectra of **2** as a function of the mixing field (0-9 mG) experienced during polarization transfer. The NMR tube was mixed with *p*-H_2 inside a voltage-controlled coil that was placed inside a mu-metal shielded chamber for these measurements.

The polarization of **2** is achieved through the creation of an initially identical AA'BC spin network on the catalyst as with **1**, where the J_{NN} is now approximately 20 Hz and $J_{(trans\text{-Hydride})N}$ = 16 Hz and $J_{(cis\text{-Hydride})N}$ <1 Hz (neglecting the $^2J_{NH}$ coupling to the α ring proton of 6.5 Hz within **2**). When **2** is bound *trans* to the NHC, the associated ^{15}N-Hydride couplings are <1 Hz. Upon ligand dissociation, the singlet state in free **2** that is created under SABRE can only evolve under the smaller weak symmetry breaking α–proton-nitrogen spin-spin couplings ($^2J_{NH}$ = 6.5 Hz, $^3J_{NH}$ < 1 Hz) and will have a longer lifetime than the bound material.

Figure 4 shows all of the detected ^{15}N resonances after SABRE transfer at 4.5 mG where additional peaks due to bound **2** within this catalyst are clearly present. A similar 'up-up-down-down' ^{15}N NMR pattern is readily seen for the two ^{15}N-coupled spins of **2** when it is located in *trans* to the NHC (δ 362.8 and 278.8) in the SABRE catalyst [Ir(H)$_2$(NHC)(**2**)$_3$]Cl, as the additional symmetry breaking couplings to hydride are now much weaker**.** The peaks with significantly reduced amplitude correspond to the more rapidly exchanging equatorial-ligand (δ 302.6 and 299.7) that couple strongly

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to hydride and consequently relax more rapidly. Confirmation of singlet character in these probed states was again provided by small tip-angle pulse examination which leads to the detection of two outer transitions in all cases (Figure 5). The process of substrate dissociation from the iridium catalyst returns to the symmetric $^{15}N_2$-environment of **2** in these measurements, as proposed earlier, and thereby promotes further, albeit slower singlet state decoherence. The observation of these signals in bound **2** is therefore reflective of indirect confirm that **2** was initially present in the singlet form.

The signal enhancements for the less sterically demanding **2** were significantly higher than those achieved for **1** under these SABRE conditions and a ^{15}N control signal (Figure 4b) confirmed the enhancement factor was now 4800 at 11.75 T (*ca*. 2%). Changing the SABRE catalyst to a *tert*-butyl-substituted catalyst[49] raised this level to 14,500-fold (~ 6%) under similar conditions. Consequently, the lifetime over which signal in the 'bound' ligand remained visible was less than 10 s in accordance with a rapid ligand loss rate of *ca*. 0.4 s^{-1} which leads to rapid cycling of this material.[38]

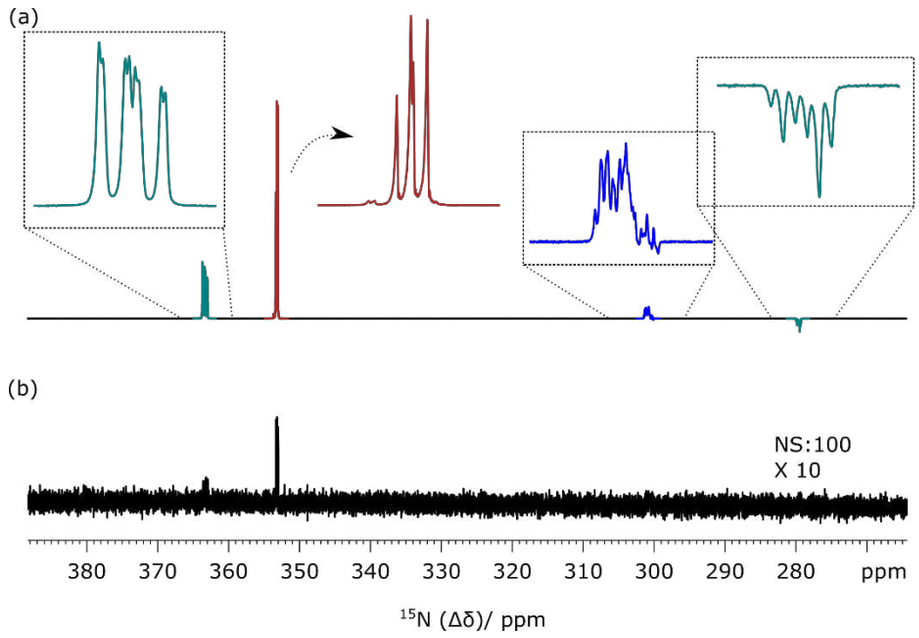

FIGURE 4: (a) High field single shot ^{15}N NMR SABRE-SHEATH spectrum of **2** after polarization transfer at 4.5 mG. Expansions show the 'free-**2**' peak at 353 ppm (red) and 'bound' axial ligand peaks (green, dominant) and the equatorial ligand signals (blue for the bound nitrogens) with characteristic singlet features. (b) ^{15}N thermal polarized NMR spectrum using 100 transients that is vertically scaled by 10 compared to (a).

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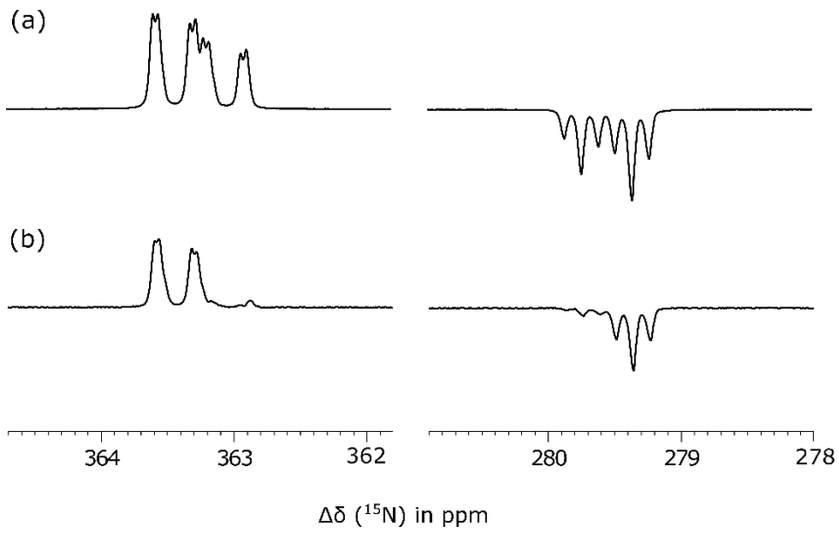

FIGURE 5: ^{15}N NMR spectra showing the axially-bound ligand peaks of **2** $[Ir(H)_2(IMes)(\mathbf{2})_3]Cl$ that are visible after SABRE-SHEATH and through (a) a 90˚ pulse and (b) a 9˚ pulse.

The lifetime of the magnetism responsible for the signal of hyperpolarized **2** was then studied in more detail. Its high-field lifetime time proved to be 5.8 ± 0.2 s. A lifetime of 21.0 ± 6.3 s was then determined for storage in the mu-metal shield, while upon storage at 0.3 T it became 56.0 ± 2.5 s. Figure 6 shows the normalized hyperpolarized signal amplitude observed for **2** in these three storage fields. Table 1 details the enhancement factor and lifetimes of **2**. These results are again affected by the catalyst and substrate concentration and when a 50-fold excess of **2** when compared to catalyst was utilized, these signal lifetimes were increased by ~40%. This scale of change is similar to that of previous reports and is a consequence of the catalysts contribution to the singlet sate decoherence being reduced, although the contribution of the intraligand H-^{15}N coupling to signal decay remains.[28, 44]

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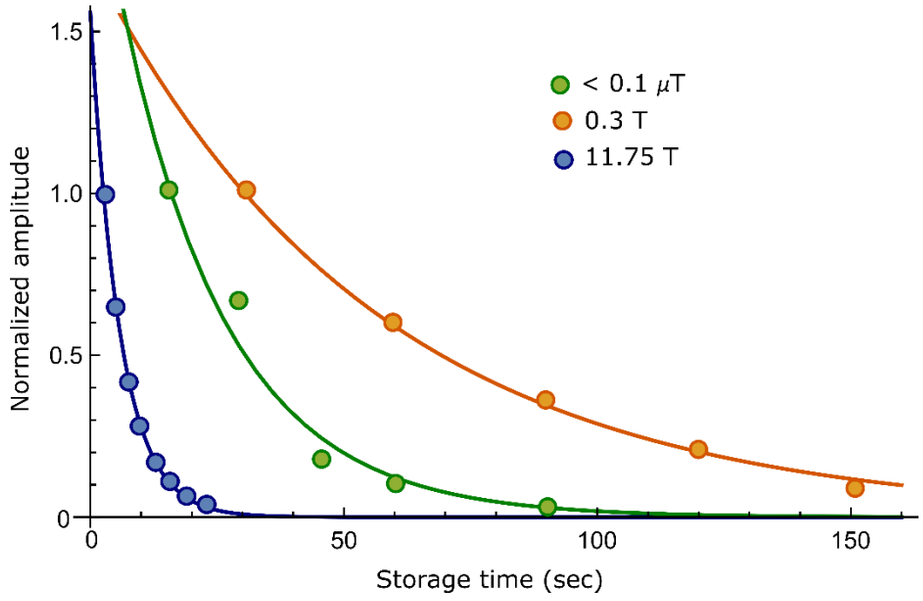

FIGURE 6: Normalized signal amplitude of ^{15}N hyperpolarized NMR signals of **2** (circles) observed from after SABRE SHEATH as a function of sample storage time. Data points were fitted to exponentials (solid curves) and results are detailed in Table 1. Three different magnetic storage fields were used: 11.75 T (blue), 0.3 T (orange), and 0 T (green).

III. CONCLUSION

In summary, we have reported how SABRE hyperpolarization can improve the ^{15}N detectability of 3-chloro-6-methoxy-4,5-d_2-pyridazine-$^{15}N_2$ (**1**) and phthlazine-$^{15}N_2$ (**2**). These molecules were synthesized as representative examples of pyridazine derivatives that possess a strong ^{15}N-^{15}N coupling (~20 Hz). Consequently, we expected to be able to prepare them in a singlet state through low-field polarization transfer via a SABRE catalyst of the form $[Ir(H)_2(NHC)(sub)_3]Cl$ where the associated hydride-hydride coupling will be of the order of -8 Hz. In the case of **1**, the steric bulk of the agent limits the efficiency of SABRE transfer such that a 0.5% polarization level is achieved, however, the isolated spin system exhibits an impressive NMR signal lifetime of 155 s when stored inside a mu metal shield. In the case of **2** it is easier to achieve higher levels of hyperpolarization due to the reduced steric bulk of this agent. Consequently, when a *tert*-butyl-substituted precatalyst is

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employed 6% ^{15}N polarization is achieved. This hyperpolarization is readily read out by breaking the symmetry of the spin system of **2** through binding to the catalyst with the result that two strong inequivalent signals are detected in the associated ^{15}N NMR responses of bound **2** when it lies *trans* to the NHC in $[Ir(H)_2(NHC)(\mathbf{2})_3]Cl$. Again, rapid ligand exchange with the SABRE catalyst reduces the apparent signal lifetime to 75 s for a 50-fold excess of reagent at a 0.3 T storage field.[47] This effect arises because ligand binding leads to a situation where $\delta\Delta$ for the two ^{15}N sites increases from 0 Hz in free **2** to ~4000 Hz when bound at 11. 75 T, depending on the ligand geometry, while introducing a further J_{HN} coupling of ~20 Hz when bound *trans* to hydride with J_{HH} = - 8 Hz and J_{NN} ~|20| Hz. These couplings and chemical shift changes enable the initially created singlet order to interconvert into the triplet manifold thereby further reducing signal lifetime. This effect is substantial, leading to a 40% fall in signal lifetime on moving from a 50-fold to 6-fold ligand excess. While we expect further catalyst optimizations to dramatically increase these levels of ^{15}N-signal gain but note it will be important to remove the catalyst if the period over which a signal is to be detected is maximized. This will be especially true if *in vivo* ^{15}N measurement is the aim.

METHODS

$^{15}N_2$-d_2-maleic hydrazide. $^{15}N_2$-Hydrazine sulfate (500 mg, 3.79 mmol, 1.0 eq) was added to a stirred solution of d_2-maleic anhydride (500 mg, 5.0 mmol, 1.32 eq) in water (7 mL). The resulting solution was heated to 100 °C for 3 h before being allowed to cool to rt. The reaction was filtered, the precipitate was collected and dried under reduced pressure to give $^{15}N_2$-d_2-maleic hydrazide as a white solid which was used in the next step without further purification.

$^{15}N_2$-3,6-dichloro-4,5-d_2-pyridazine.$^{15}N_2$-d_2-Maleic hydrazide (325 mg, 2.80 mmol, 1.0 eq.) in $POCl_3$ (3.0 mL) was heated to 95 °C for 3 h. Then the reaction was cooled to rt and added dropwise to an ice cold solution of $NaHCO_3$ to neutralise. EtOAc (15 mL) was added and the two layers were separated. The aqueous layer was extracted with EtOAc (3 x 15 mL) and the combined organic layers were dried ($MgSO_4$) and concentrated under reduced pressure to give the crude product.

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Purification by flash column chromatography with 8:2 hexane-EtOAc as eluent gave $^{15}N_2$-3,6-dichloro-4,5-d_2-pyridazine (321 mg, 75%) as a white solid, R_F (8:2 hexane-EtOAc) 0.3; **^{13}C NMR** (126 MHz, $CDCl_3$) δ (ppm) 156.0 (t, J = 7.24 Hz), 130.0 (t, J = 26.6 Hz); **^{15}N NMR** (51 MHz, $CDCl_3$) δ (ppm) 390.2 (s); **MS** (ESI) *m/z* 175 [$(M + Na)^+$, 40] 153 [$(M + H)^+$, 100]; **HRMS** (ESI) *m/z* $[M + Na]^+$ calculated for $C_4Cl_2D_2{}^{15}N_2$ 174.9553, found 174.9559 (−3.0 ppm error).

$^{15}N_2$-3-chloro- 4,5-d_2-6-methoxypyridazine (1). Sodium methoxide (60 mg, 1.1 mmol, 1.1 eq.) was added to a stirred solution of $^{15}N_2$-3,6-dichloro-4,5-d_2-pyridazine (153 mg, 1.0 mmol, 1.0 eq) in MeOH (10 mL) and the resulting solution was stirred at rt for 48 h. The reaction was concentrated under reduced pressure to give the crude product. Purification by flash column chromatography with 95:5-85:15 CH_2Cl_2-EtOAc as eluent gave **1** (143 mg, 97%) as a white solid, R_F (85:15 CH_2Cl_2-EtOAc) 0.3; **^{1}H NMR** (500 MHz, $CDCl_3$) 4.12 (s, 3H); **^{13}C NMR** (126 MHz, $CDCl_3$) δ (ppm) 164.4 (d, J = 5.3 Hz), 151.0 (m), 130.3 (app. t, J = 24.3 Hz), 119.7 (dd, J = 23.4, 3.8 Hz), 55.2 (d, J = 4.0 Hz); **^{15}N NMR** (41 MHz, $CDCl_3$) δ (ppm) 372.3 (d, J = 23.7 Hz), 339.9 (d, J = 23.7 Hz); **MS** (ESI) *m/z* 171 [$(M + Na)^+$, 80] 149 [$(M + H)^+$, 100]; **HRMS** (ESI) *m/z* $[M + Na]^+$ calculated for $C_5H_3ClD_2{}^{15}N_2O$ 171.0049, found 171.0053 (−1.7 ppm error).

$^{15}N_2$-phthalazine(2). A solution of $^{15}N_2H_4.H_2SO_4$ (1.21 g, 9.31 mmol) in 1 M NaOH (15 mL) was added to a solution of phthaldialdehyde (1.25 g, 9.33 mmol) and EtOH (30 mL) at room temperature, and stirred for 3 hours. The resulting solution was extracted with DCM (3 × 100 mL) and the combined extracts concentrated *in vacuo*. Purification by column chromatography (EtOAc) afforded **2** (815 mg, 66%) as an orange solid. **^{1}H NMR** (400 MHz, $CDCl_3$) δ (ppm) 9.43 (app t, J = 8.2 Hz, 2H) 7.87-7.81 (m, 4H); **^{13}C NMR** (101 MHz, $CDCl_3$) δ (ppm) 151.1 (t, J = 4.4 Hz), 132.7, 126.4 (t, J =1.8 Hz), 126.2; **^{15}N NMR** (51 MHz, $CDCl_3$) δ (ppm) 365.3; **MS** (ESI) *m/z* 155 [$(M + Na)^+$, 100], 133 [$(M + H)^+$, 80]; **HRMS** (ESI) *m/z* $[M + H]^+$ calculated for $C_8H_7{}^{15}N_2$ 133.0544, found 133.0548 (−2.5 ppm error).

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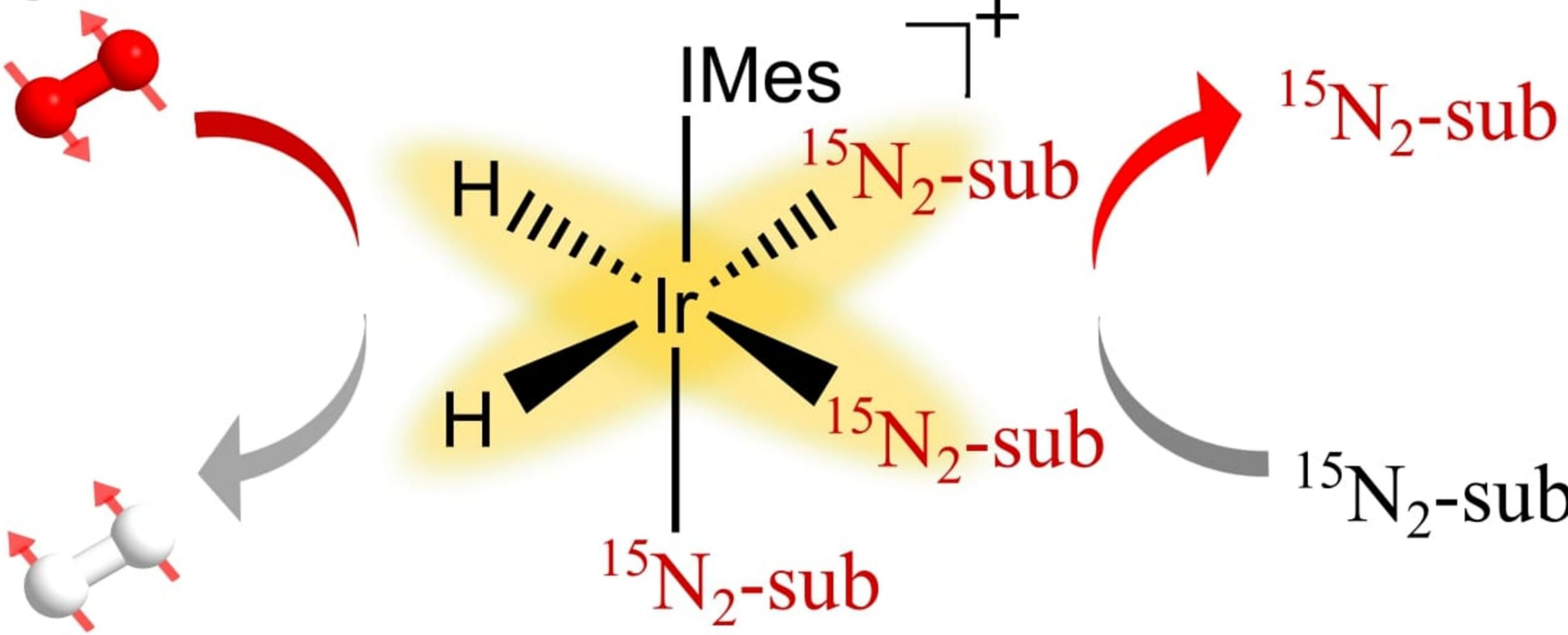
(a)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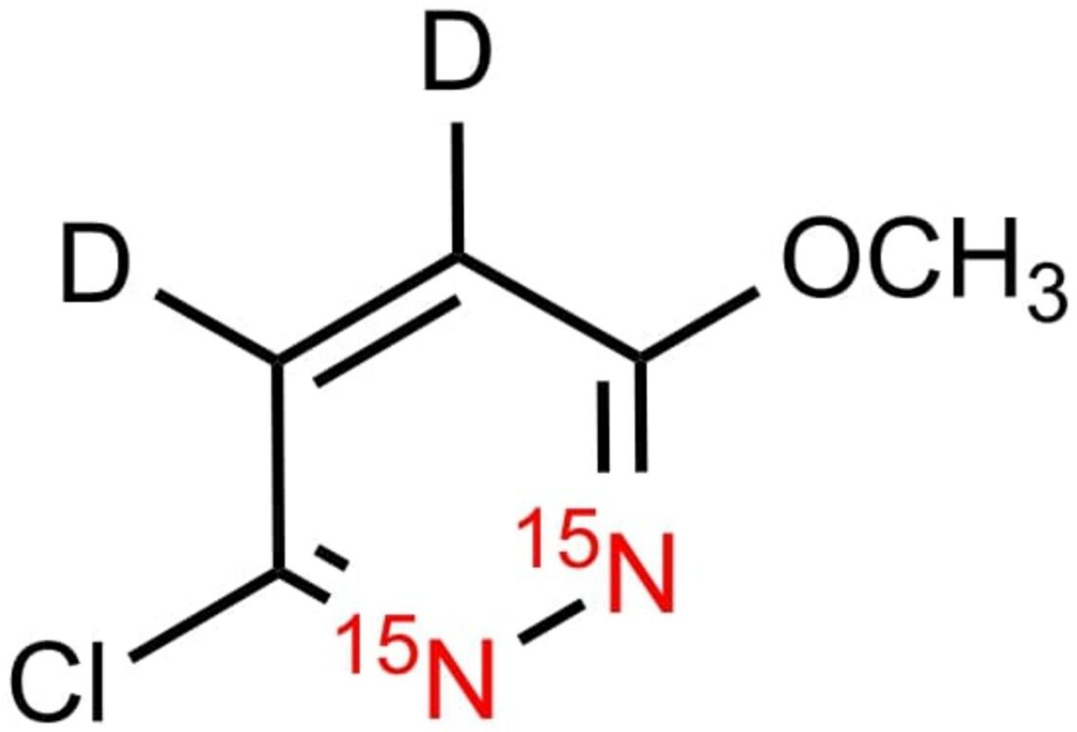
(b)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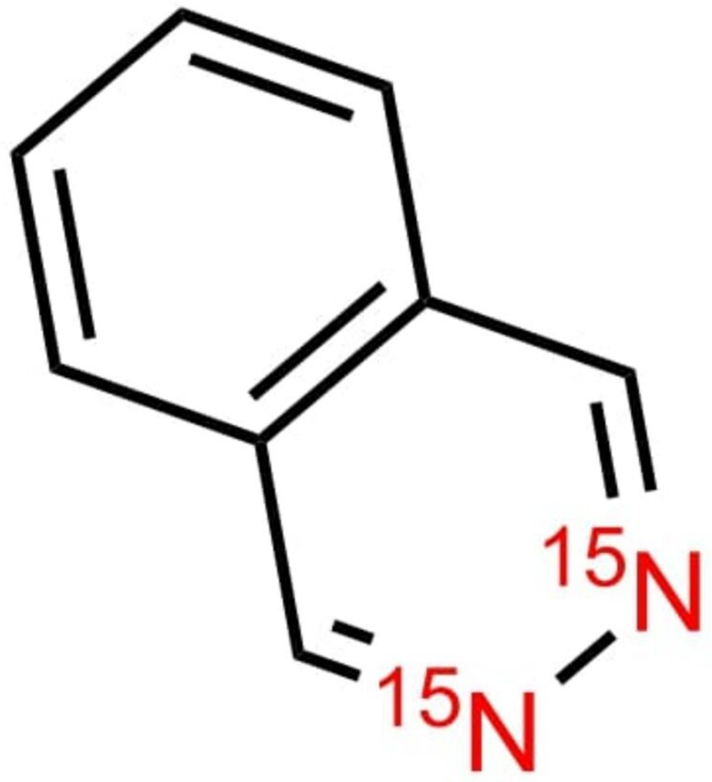
Cl
15N
15N

1


15N
15N

2

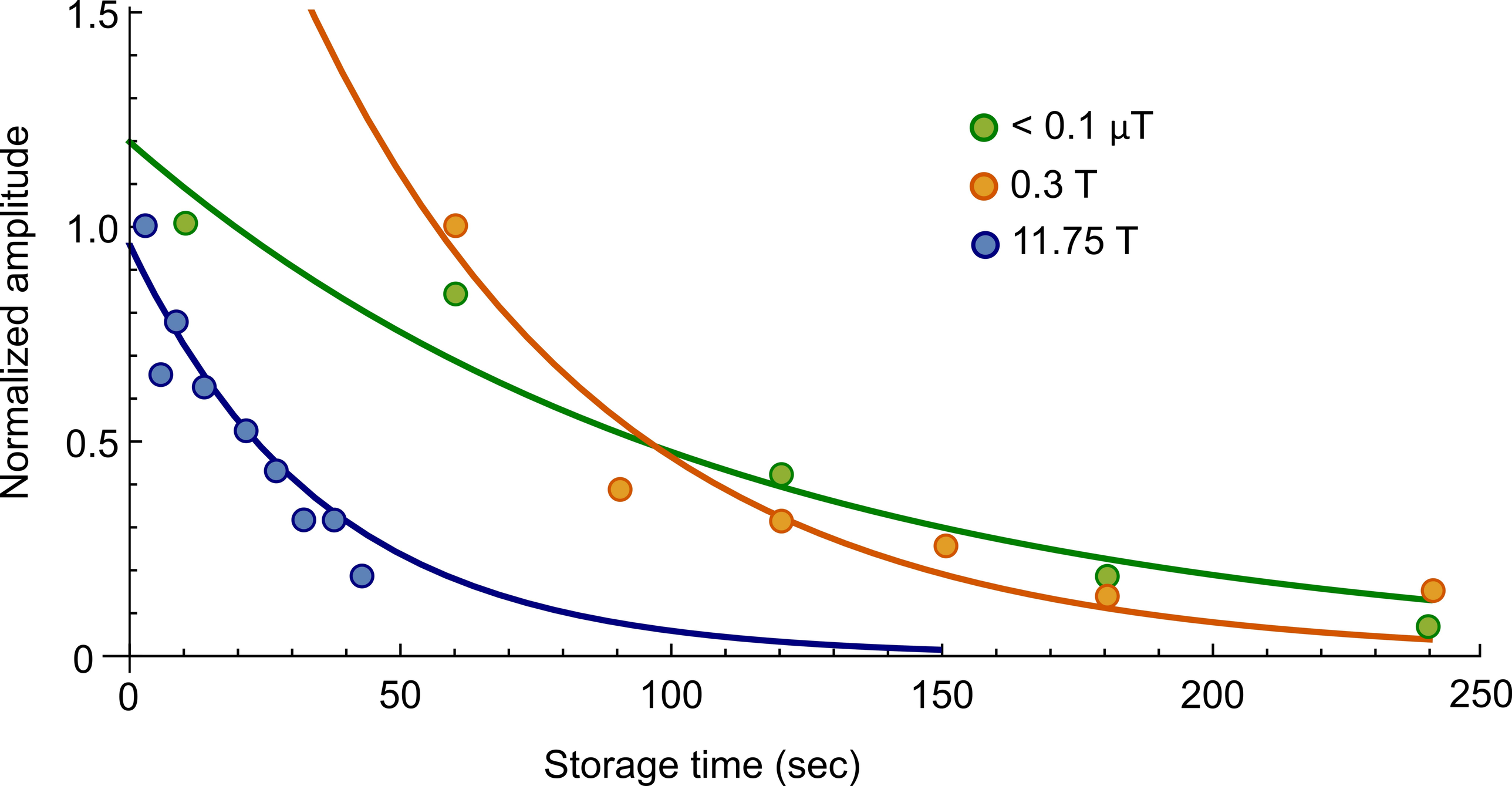
(a)

(b)

(c)

370 360 350 340 330

$\Delta\delta(^{15}N)$/ppm


Normalized amplitude
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0.5
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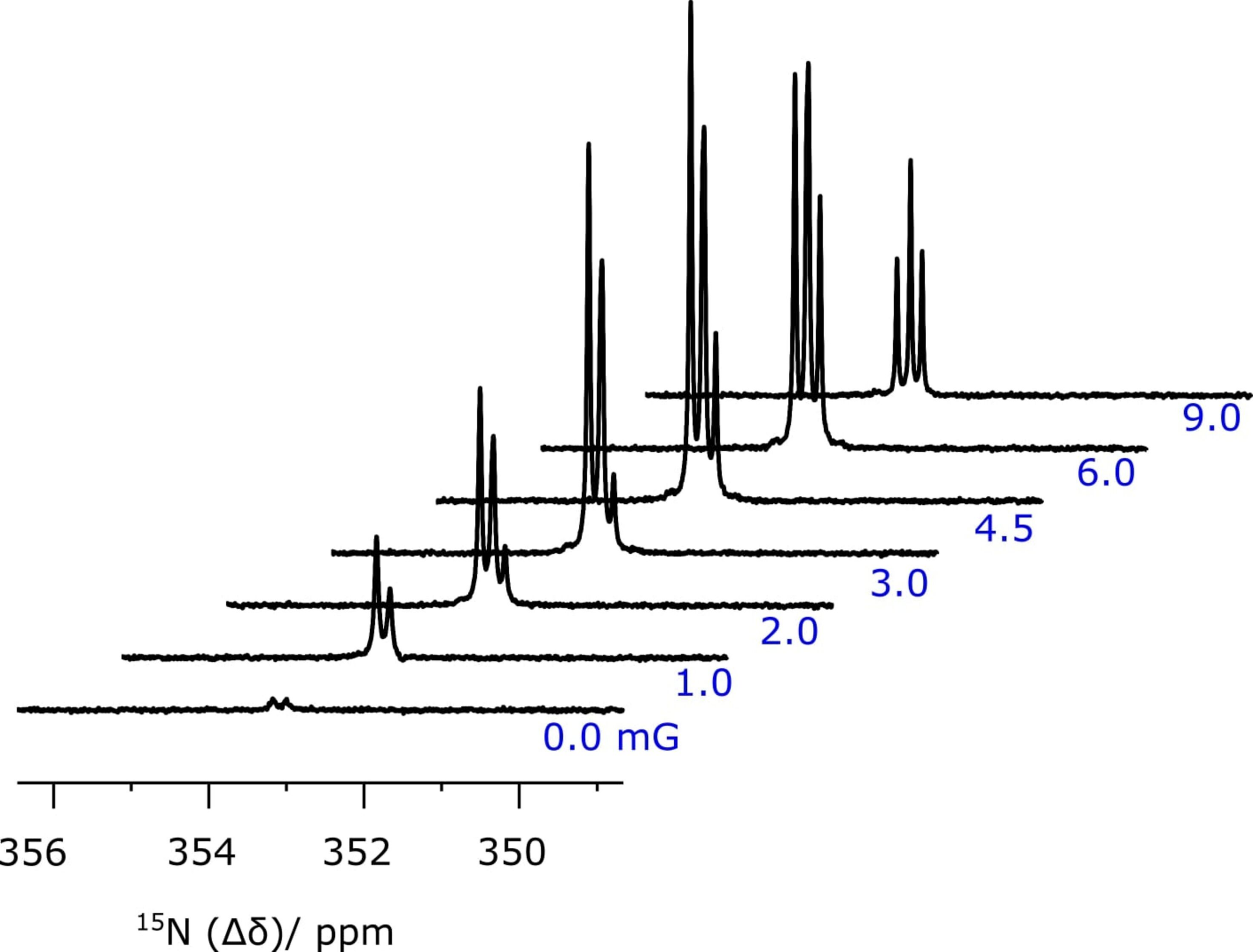
0
50
100
150
200
250
Storage time (sec)
< 0.1 μT
0.3 T
11.75 T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6.0
4.5
3.0
2.0
1.0
0.0 mG
356
354
352
350
15N (Δδ)/ ppm

(a)

(b)

NS:100
X 10

380 370 360 350 340 330 320 310 300 290 280 ppm

^{15}N (Δδ)/ ppm

(a)

(b)

364 363 362

280 279 278

Δδ (^{15}N) in ppm

Normalized amplitude

1.5
1.0
0.5
0

0
50
100
150

Storage time (sec)

< 0.1 μT
0.3 T
11.75 T